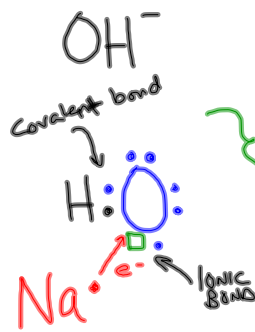
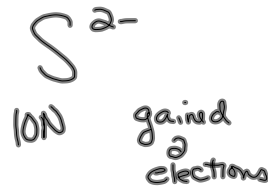
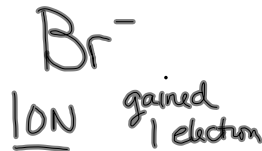
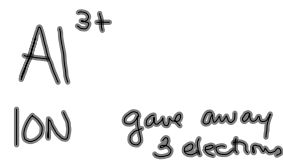
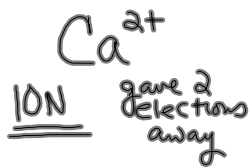
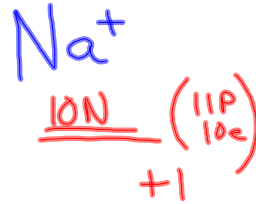
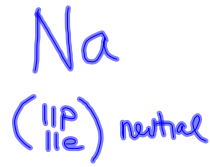
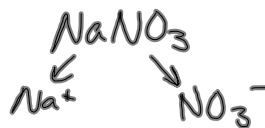
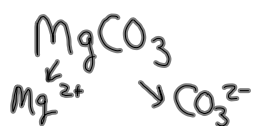


An ion is an atom
(or group of atoms) with
a charge.

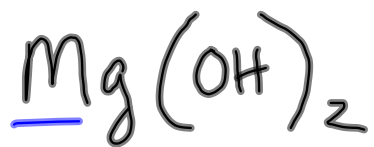


the O and H are covalently bonded - this is a strong bond - we refer to this as an "OH group" - the OH group alone still needs 1 more v.e. So... an atom of an element in column 1 could give up a v.e.

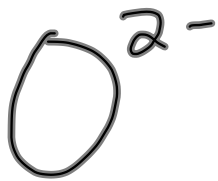
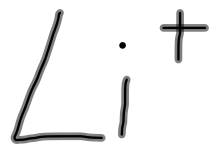
creating an Na⁺ ion and... an OH⁻ ion



giving
up
2ve



Ions come from ionic bonds -
where atoms gain + lose
electrons!



"oxide"



"bromide"

CO_2 — Covalent bond

